**General Final Review – Equations**

For each of the following reactions, do the following:

* Indicate the type of reaction
* Indicate whether the reaction will occur (using solubility tables and the activity series).
* If the reaction does occur, determine what products will be formed.

\_\_\_\_\_\_\_\_ Na + \_\_\_\_\_\_\_\_ Pb(OH)2 →

\_\_\_\_\_\_\_\_ Ca(OH)2 + \_\_\_\_\_\_\_\_ HBr →

\_\_\_\_\_\_\_\_ Pb(NO3)2 + \_\_\_\_\_\_\_\_ NaI →

\_\_\_\_\_\_\_\_ KNO3 + \_\_\_\_\_\_\_\_ Mg(C2H3O2)2 →

\_\_\_\_\_\_\_\_ KBr + \_\_\_\_\_\_\_\_ ZnF2 →

Write the complete balanced equation for the combustion of C2H2. (Show symbols of state, other stuff like that)

Write the complete balanced equation for the following reaction:

* When potassium metal is placed in liquid water, a big explosion takes place. This explosion creates powdered potassium hydroxide and hydrogen gas.